

Modern Chemistry Chemical Equilibrium Answer Key

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K_p (equilibrium constant) is independent of pressure and concentration. 34 .One mole of a compound AB reacts with one mole of a compound CD according to the equation $AB + CD \rightleftharpoons AD + CB$. When equilibrium had been established it was found that 34 mole each of reactant AB and CD had been converted to AD and CB.

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Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship

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between the ion product and the K_{sp} for that solution, listed on the left. ____ The ion product exceeds the K_{sp} .

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14) Gas X_2 reacts with gas Y_2 to give gaseous XY according to the equation $X_2 + Y_2 \rightleftharpoons 2XY$. $K_{eq} = 2.6 \times 10^{-3}$. What is K_{eq} for the reverse reaction? a) 6.8×10^{-6} b) 5.1×10^{-2} c) 2.0×10^1 d) 3.8×10^2 15) Which one of the following statements about a reaction indicates that the products are favored at equilibrium?

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An equilibrium mixture at a specific temperature is found to consist of $1.2 \times 10^{-3} \text{ mol} / \text{LHCl}$, $3.8 \times 10^{-4} \text{ mol} / \text{L}$, $5.8 \times 10^{-2} \text{ mol} / \text{L}$, $5.8 \times 10^{-2} \text{ mol} / \text{L}$, and $5.8 \times 10^{-2} \text{ mol} / \text{L}$ according to the following:

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When two volumes of gas are in equilibrium, if one volume increases with increased pressure, the other volume must decrease with decreased pressure. When two volumes of gas are in equilibrium, if...

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